N25 – Gases Review

KMT = Kinetic Molecular Theory

1. Gases consist of large numbers of tiny particles that are far apart relative to their size
   – the volume of each gas molecule is considered negligible, they are treated as point particles.
2. Gases are "Ideal Gases" – meaning they do not interact with each other.
   – There are no forces of attraction or repulsion between particles.
3. Gas particles undergo elastic collisions
   – meaning they do not lose energy when colliding.
4. Gas particles are in a constant, rapid, straight line motion
   – they possess kinetic energy (motion energy).
5. The average kinetic energy of the particles is proportional to temperature
   – (in Kelvin!!!) – T ↑, KE ↑
6. There is a distribution of speeds, some go faster than others
   – so overall there is an average kinetic energy of the sample.

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